

The University of the State of New York
REGENTS HIGH SCHOOL EXAMINATION

**PHYSICAL SETTING
CHEMISTRY**

Thursday, August 17, 2023 — 8:30 to 11:30 a.m., only

The possession or use of any communications device is strictly prohibited when taking this examination. If you have or use any communications device, no matter how briefly, your examination will be invalidated and no score will be calculated for you.

This is a test of your knowledge of chemistry. Use that knowledge to answer all questions in this examination. Some questions may require the use of the *2011 Edition Reference Tables for Physical Setting/Chemistry*. You are to answer *all* questions in all parts of this examination according to the directions provided in this examination booklet.

A separate answer sheet for Part A and Part B–1 has been provided to you. Follow the instructions from the proctor for completing the student information on your answer sheet. Record your answers to the Part A and Part B–1 multiple-choice questions on this separate answer sheet. Record your answers for the questions in Part B–2 and Part C in your separate answer booklet. Be sure to fill in the heading on the front of your answer booklet.

All answers in your answer booklet should be written in pen, except for graphs and drawings, which should be done in pencil. You may use scrap paper to work out the answers to the questions, but be sure to record all your answers on your separate answer sheet or in your answer booklet as directed.

When you have completed the examination, you must sign the statement printed on your separate answer sheet, indicating that you had no unlawful knowledge of the questions or answers prior to the examination and that you have neither given nor received assistance in answering any of the questions during the examination. Your answer sheet and answer booklet cannot be accepted if you fail to sign this declaration.

Notice . . .

A four-function or scientific calculator and a copy of the *2011 Edition Reference Tables for Physical Setting/Chemistry* must be available for you to use while taking this examination.

DO NOT OPEN THIS EXAMINATION BOOKLET UNTIL THE SIGNAL IS GIVEN.

Part A

Answer all questions in this part.

Directions (1–30): For *each* statement or question, record on your separate answer sheet the *number* of the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the *2011 Edition Reference Tables for Physical Setting/Chemistry*.

- 1 Which proposal in the development of the modern model of the atom was made before the others?
- (1) Atoms are hard, indivisible spheres of different sizes.
 - (2) Atoms are mostly empty space with a small dense nucleus.
 - (3) Atoms have electrons that have wavelike properties.
 - (4) Atoms have an internal structure that contains negative particles.
- 2 According to the wave-mechanical model, in the ground state, the 10 electrons of a neon atom would be located
- (1) in the nucleus
 - (2) in orbitals
 - (3) in the first shell
 - (4) in the valence shell
- 3 Which statement describes two different isotopes of carbon?
- (1) The isotopes contain the same number of neutrons and have the same atomic number.
 - (2) The isotopes contain the same number of neutrons but have a different atomic number.
 - (3) The isotopes contain a different number of neutrons but have the same atomic number.
 - (4) The isotopes contain a different number of neutrons and have a different atomic number.
- 4 An element that is a very reactive metal could have an atomic number of
- (1) 9
 - (2) 2
 - (3) 19
 - (4) 79
- 5 Which element is listed with the number of protons in each of its atoms?
- (1) nitrogen, 14
 - (2) silicon, 14
 - (3) oxygen, 16
 - (4) phosphorus, 16
- 6 What is the overall charge on the nucleus of a fluorine atom?
- (1) -1
 - (2) -9
 - (3) $+9$
 - (4) $+19$
- 7 As the elements with atomic numbers 11 through 17 are considered in order of increasing atomic number, the classification of the elements changes from
- (1) metal to metalloid to nonmetal
 - (2) metal to nonmetal to metalloid
 - (3) nonmetal to metalloid to metal
 - (4) nonmetal to metal to metalloid
- 8 At STP, which property can be used to differentiate one-mole samples of $\text{Cl}_2(\text{g})$ and $\text{Kr}(\text{g})$?
- (1) phase
 - (2) pressure
 - (3) chemical reactivity
 - (4) temperature
- 9 Two gaseous forms of oxygen are diatomic oxygen, O_2 , and ozone, O_3 . These two forms of oxygen have
- (1) the same molecular structure and the same properties
 - (2) the same molecular structure and different properties
 - (3) different molecular structures and the same properties
 - (4) different molecular structures and different properties
- 10 A compound consists of two or more different elements that are
- (1) physically mixed in a fixed proportion
 - (2) physically mixed in a proportion that can vary
 - (3) chemically combined in a fixed proportion
 - (4) chemically combined in a proportion that can vary

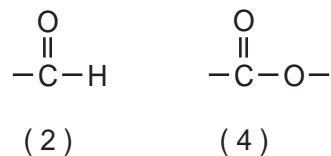
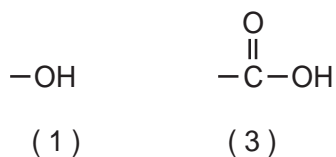
- 11 Compared to the chemical and physical properties of the compound CO, the compound CO₂ has
- (1) the same chemical properties and the same physical properties
 - (2) the same chemical properties and different physical properties
 - (3) different chemical properties and the same physical properties
 - (4) different chemical properties and different physical properties
- 12 Which phrase describes the molecular polarity of and the charge distribution in an HCl molecule?
- (1) nonpolar with an asymmetrical charge distribution
 - (2) nonpolar with a symmetrical charge distribution
 - (3) polar with an asymmetrical charge distribution
 - (4) polar with a symmetrical charge distribution
- 13 Which atom forms an ion with a radius larger than the atomic radius?
- (1) calcium atom
 - (2) oxygen atom
 - (3) lead atom
 - (4) tin atom
- 14 Which sample can be classified as a substance?
- (1) air
 - (2) argon
 - (3) soil
 - (4) seawater
- 15 A 1.0-gram sample of NaCl(s) is dissolved in 100. grams of water at 25°C, and another 1.0-gram sample of NaCl(s) is dissolved in 50. grams of water at 25°C. Which property of the two resulting mixtures will be different?
- (1) color of the components in the mixture
 - (2) particle size of the components in the mixture
 - (3) polarity of the components in the mixture
 - (4) proportion by mass of the components in the mixture
- 16 Based on Table G, which compound has the greatest solubility in 100. grams of water at 10.°C?
- (1) HCl
 - (2) NaCl
 - (3) KCl
 - (4) NH₄Cl
- 17 According to the kinetic molecular theory, ideal gas particles
- (1) are separated by small distances relative to their size
 - (2) do not transfer energy when they collide with each other
 - (3) have attractive forces between them
 - (4) move in random, constant, straight-line motion
- 18 A reaction is most likely to occur when the reacting particles collide with proper orientation and proper
- (1) charge
 - (2) energy
 - (3) mass
 - (4) volume
- 19 A sample of SO₂(g) and a sample of NO₂(g) contain the same number of molecules when they have the same
- (1) mass, temperature, and volume
 - (2) mass, temperature, and pressure
 - (3) pressure, volume, and temperature
 - (4) pressure, volume, and mass
- 20 Which term identifies the strong intermolecular forces found in a sample of liquid water?
- (1) ionic bonding
 - (2) covalent bonding
 - (3) hydrogen bonding
 - (4) metallic bonding
- 21 Which mathematical expression represents the heat of reaction for a chemical change?
- (1) (PE of the products) – (PE of the reactants)
 - (2) (PE of the products) + (PE of the reactants)
 - (3) (PE of the products) ÷ (PE of the reactants)
 - (4) (PE of the products) × (PE of the reactants)

- 22 Catalysts can increase the rate of a chemical reaction by providing
- (1) an alternate reaction pathway with a higher activation energy
 - (2) the same reaction pathway with a higher activation energy
 - (3) an alternate reaction pathway with a lower activation energy
 - (4) the same reaction pathway with a lower activation energy

- 23 Which term represents the disorder of a system?
- (1) entropy
 - (2) mole
 - (3) quanta
 - (4) pressure

- 24 Which element must be present in all organic compounds?
- (1) nitrogen
 - (2) hydrogen
 - (3) carbon
 - (4) sulfur

- 25 Based on Table R, which functional group allows a compound to be classified as an organic acid?



- 26 In a voltaic cell, oxidation occurs
- (1) at the cathode
 - (2) at the anode
 - (3) in the salt bridge
 - (4) in the external circuit

- 27 Which positive ion must be present in an aqueous solution of an Arrhenius acid?
- (1) H_3O^+
 - (2) Na^+
 - (3) NH_4^+
 - (4) Rb^+

- 28 A change that converts an atom of one element to another element is called
- (1) neutralization
 - (2) oxidation
 - (3) sublimation
 - (4) transmutation

- 29 Which radiation has the *least* ionizing power and greatest penetrating power?
- (1) alpha particles
 - (2) beta particles
 - (3) gamma emissions
 - (4) positron emissions

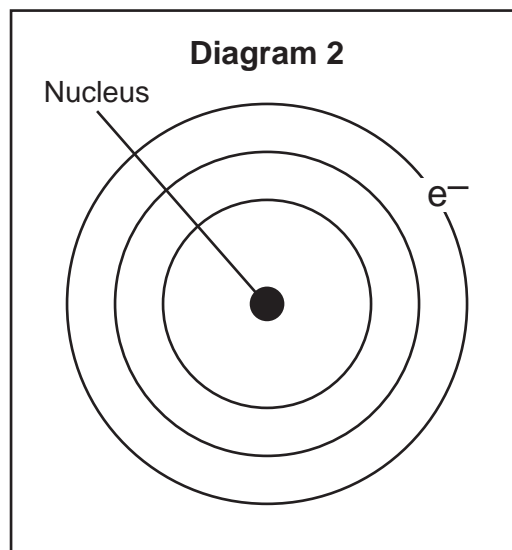
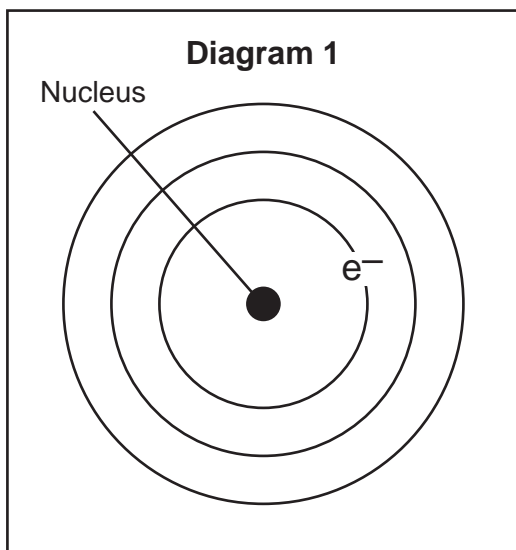
- 30 Which nuclear emission is listed with its notation?
- (1) alpha particle, ${}^4_2\text{p}$
 - (2) beta particle, ${}^0_{-1}\text{e}$
 - (3) neutron, ${}^1_0\text{n}$
 - (4) proton, ${}^1_1\text{p}$

Part B-1

Answer all questions in this part.

Directions (31–50): For each statement or question, record on your separate answer sheet the number of the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2011 Edition Reference Tables for Physical Setting/Chemistry.

31 Diagram 1 represents an atom of hydrogen, showing the nucleus and the electron in the first shell.



What is represented by diagram 2?

- (1) a positive ion of hydrogen
- (2) a negative ion of hydrogen
- (3) an atom of hydrogen in an excited state
- (4) an atom of hydrogen in the ground state

32 The table below gives the atomic mass and the abundance of the two naturally occurring isotopes of copper.

Naturally Occurring Isotopes of Copper

Isotope Notation	Atomic Mass of the Isotope (u)	Natural Abundance (%)
Cu-63	62.93	69.15
Cu-65	64.93	30.85

Which numerical setup can be used to calculate the atomic mass of the element copper?

- (1) $(62.93 \text{ u})(30.85) + (64.93 \text{ u})(69.15)$
- (2) $(62.93 \text{ u})(69.15) + (64.93 \text{ u})(30.85)$
- (3) $(62.93 \text{ u})(0.3085) + (64.93 \text{ u})(0.6915)$
- (4) $(62.93 \text{ u})(0.6915) + (64.93 \text{ u})(0.3085)$

33 Which subatomic particles were discovered as the result of experiments with cathode ray tubes?

- (1) electrons (3) positrons
(2) neutrons (4) protons

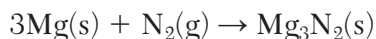
34 Which change in electron location in an atom of calcium is accompanied by the greatest amount of energy emitted?

- (1) from shell 1 to shell 2
(2) from shell 2 to shell 1
(3) from shell 1 to shell 4
(4) from shell 4 to shell 1

35 As the atomic number increases in Group 2 from Be to Ba, the first ionization energy

- (1) decreases, and the atomic radius decreases
(2) decreases, and the atomic radius increases
(3) increases, and the atomic radius decreases
(4) increases, and the atomic radius increases

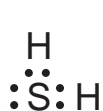
36 Given the equation representing a reaction:



What is the mass of Mg_3N_2 that is produced when 14.58 grams of magnesium completely reacts with 5.60 grams of nitrogen?

- (1) 8.98 g (3) 20.18 g
(2) 10.46 g (4) 49.34 g

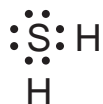
37 Which Lewis electron-dot diagram represents a molecule of H_2S ?



(1)



(3)



(2)



(4)

38 What is the vapor pressure of water at $90.^\circ\text{C}$?

- (1) 40. kPa (3) 94 kPa
(2) 68 kPa (4) 150. kPa

39 A sample is composed of two different substances that can be separated by using filter paper. This sample is classified as

- (1) a heterogeneous mixture
(2) a homogeneous mixture
(3) a single compound
(4) an unsaturated solution

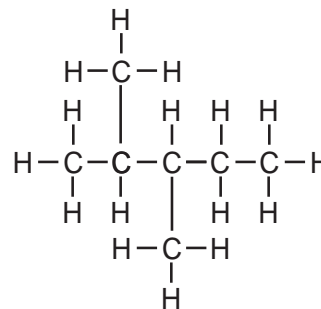
40 Which solutions react to produce an insoluble compound?

- (1) $\text{KCl}(aq) + \text{LiCl}(aq) \rightarrow$
(2) $\text{LiCl}(aq) + \text{NaNO}_3(aq) \rightarrow$
(3) $\text{NaCl}(aq) + \text{AgClO}_3(aq) \rightarrow$
(4) $\text{KNO}_3(aq) + \text{AgClO}_3(aq) \rightarrow$

41 What is the amount of heat energy absorbed when 40.0 grams of water is heated from 10.0°C to 30.0°C ?

- (1) $1.67 \times 10^3 \text{ J}$ (3) $5.02 \times 10^3 \text{ J}$
(2) $3.34 \times 10^3 \text{ J}$ (4) $2.67 \times 10^5 \text{ J}$

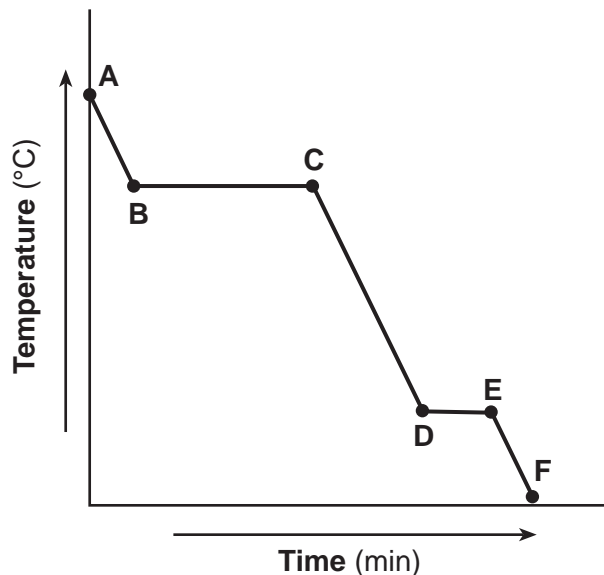
42 Given the formula representing a compound:



What is a chemical name of this compound?

- (1) 2,3,3-trimethylbutane
(2) 2-methyl-2-ethylbutane
(3) 2,3-dimethylpentane
(4) 2,3-ethylpentane

43 Given the cooling curve for a substance:



During which intervals is the potential energy decreasing while the average kinetic energy remains constant?

- (1) *BC* and *AB* (3) *EF* and *AB*
(2) *BC* and *DE* (4) *EF* and *DE*
- 44 Which formula represents a hydrocarbon?
(1) CO_2 (3) $\text{C}_2\text{H}_4\text{O}_2$
(2) C_3H_8 (4) $\text{C}_3\text{H}_8\text{O}$
- 45 Which type of organic reaction produces carbon dioxide and water?
(1) combustion (3) fermentation
(2) esterification (4) saponification

46 Which half-reaction equation represents the oxidation of lithium?

- (1) $\text{Li}^+ + \text{e}^- \rightarrow \text{Li}$ (3) $\text{Li} \rightarrow \text{Li} + \text{e}^-$
(2) $\text{Li} + \text{e}^- \rightarrow \text{Li}^+$ (4) $\text{Li} \rightarrow \text{Li}^+ + \text{e}^-$
- 47 Based on Table *J*, which ionic equation represents a spontaneous reaction that can occur in a voltaic cell?
(1) $\text{Fe}(\text{s}) + \text{Mg}^{2+}(\text{aq}) \rightarrow \text{Fe}^{2+}(\text{aq}) + \text{Mg}(\text{s})$
(2) $\text{Fe}(\text{s}) + \text{Mg}(\text{s}) \rightarrow \text{Fe}^{2+}(\text{aq}) + \text{Mg}^{2+}(\text{aq})$
(3) $\text{Fe}^{2+}(\text{aq}) + \text{Mg}^{2+}(\text{aq}) \rightarrow \text{Fe}(\text{s}) + \text{Mg}(\text{s})$
(4) $\text{Fe}^{2+}(\text{aq}) + \text{Mg}(\text{s}) \rightarrow \text{Fe}(\text{s}) + \text{Mg}^{2+}(\text{aq})$
- 48 Given the equation representing a system at equilibrium:
$$\text{NH}_3(\text{g}) + \text{H}_2\text{O}(\ell) \rightleftharpoons \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq})$$

In this system, the $\text{H}_2\text{O}(\ell)$ acts as
(1) an acid, because it accepts an H^+
(2) an acid, because it donates an H^+
(3) a base, because it accepts an H^+
(4) a base, because it donates an H^+
- 49 Which indicator is blue in a solution that has a pH value of 7.0?
(1) bromocresol green (3) phenolphthalein
(2) methyl orange (4) thymol blue
- 50 Which radioisotope requires long-term storage to prevent the risk of biological exposure?
(1) N-16 (3) Pu-239
(2) K-42 (4) Au-198

Part B–2

Answer all questions in this part.

Directions (51-65): Record your answers in the spaces provided in your answer booklet. Some questions may require the use of the *2011 Edition Reference Tables for Physical Setting/Chemistry*.

Base your answers to questions 51 and 52 on the information below and on your knowledge of chemistry.

The four naturally occurring isotopes of sulfur are sulfur-32, sulfur-33, sulfur-34, and sulfur-36.

51 State the number of valence electrons in an atom of sulfur-32 in the ground state. [1]

52 State the number of neutrons in an atom of sulfur-33. [1]

Base your answers to questions 53 through 55 on the information below and on your knowledge of chemistry.

The Group 1 elements below hydrogen are called alkali metals, and the first five Group 17 elements are called halogens.

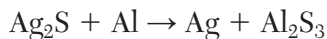
53 Compare the electrical conductivity of an alkali metal at STP with the electrical conductivity of a halogen at STP. [1]

54 State the type of chemical bonds formed when an alkali metal and a halogen react to form a compound. [1]

55 Explain, in terms of electrons, why the halogens are all placed in the same group on the Periodic Table. [1]

Base your answers to questions 56 through 58 on the information below and on your knowledge of chemistry.

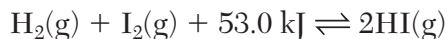
The unbalanced equation below represents the reaction between silver sulfide and aluminum.



- 56 Balance the equation *in your answer booklet* for the reaction, using the *smallest* whole number coefficients. [1]
- 57 Determine the gram-formula mass of the aluminum sulfide product. [1]
- 58 Show a numerical setup for calculating the number of moles of silver sulfide in a 546-gram sample of Ag_2S (gram-formula mass = 248 g/mol). [1]
-

Base your answers to questions 59 through 61 on the information below and on your knowledge of chemistry.

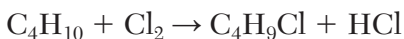
Hydrogen gas and iodine gas can combine in a reversible reaction to form hydrogen iodide gas. The equation below represents this system at equilibrium in a sealed, rigid container.



- 59 State evidence from the equation that the forward reaction is endothermic. [1]
- 60 State, in terms of reaction rates, why the concentration of $\text{HI}(\text{g})$ remains constant when the system is at equilibrium. [1]
- 61 State how an increase in temperature affects the concentration of $\text{HI}(\text{g})$. [1]
-

Base your answers to questions 62 through 65 on the information below and on your knowledge of chemistry.

The equation below represents the reaction of straight-chain butane with chlorine.



The reaction produces two different isomers of $\text{C}_4\text{H}_9\text{Cl}$: 1-chlorobutane and 2-chlorobutane. The table below shows the boiling point of each $\text{C}_4\text{H}_9\text{Cl}$ isomer at standard pressure.

Name of Compound	Boiling Point (°C)
1-chlorobutane	78.4
2-chlorobutane	68.2

- 62 State, in terms of carbon-carbon bonds, why butane is saturated. [1]
- 63 Identify the class of organic compounds to which the organic product in the equation belongs. [1]
- 64 Based on the boiling point data, compare the strength of the intermolecular forces in 1-chlorobutane to the strength of the intermolecular forces in 2-chlorobutane. [1]
- 65 Draw a structural formula for the 2-chlorobutane. [1]
-

Part C

Answer all questions in this part.

Directions (66-85): Record your answers in the spaces provided in your answer booklet. Some questions may require the use of the *2011 Edition Reference Tables for Physical Setting/Chemistry*.

Base your answers to questions 66 through 68 on the information below and on your knowledge of chemistry.

During a laboratory activity, appropriate safety equipment was used and safety procedures were followed, as two students worked with hydrated CuSO_4 . A hydrate is a compound that has water within its crystal structure.

Student 1 and student 2 each heated their own 3.00-gram sample of hydrated CuSO_4 . Student 2 determined the percent composition by mass of water in the hydrated CuSO_4 to be 37.0%. The accepted value for the percent of water in the hydrated CuSO_4 is 36.1%. The final mass values for each student sample are shown in the table below.

Mass Values for Lab Activity

Student	Mass of Hydrated CuSO_4 (g)	Mass of CuSO_4 After Final Heating (g)	Mass of H_2O in Hydrated CuSO_4 (g)
1	3.00	1.98	1.02
2	3.00	1.89	1.11

- 66 State the number of significant figures in the mass of the hydrated CuSO_4 sample used by student 1. [1]
- 67 Determine the percent composition by mass of water in the hydrated CuSO_4 sample, based on the data for student 1. [1]
- 68 Show a numerical setup for calculating the percent error for the percent composition by mass of water in the hydrated CuSO_4 sample as determined by student 2. [1]
-

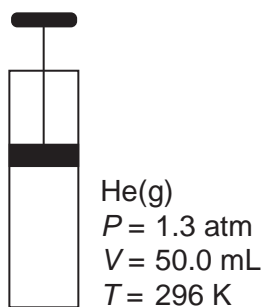
Base your answers to questions 69 through 71 on the information below and on your knowledge of chemistry.

Ammonia, $\text{NH}_3(\text{g})$, is a compound that is used to manufacture fertilizer for agriculture. This gas is produced by a reaction between nitrogen gas, $\text{N}_2(\text{g})$, and hydrogen gas, $\text{H}_2(\text{g})$. During this reaction, bonds are both broken and formed. Bond energy is the amount of energy required to break one mole of bonds between atoms in the gas phase. The bond energy of an H–H bond is 436 kilojoules per mole.

- 69 State the number of electrons shared between the nitrogen atom and one of the hydrogen atoms in an ammonia molecule. [1]
- 70 Identify the noble gas that has atoms in the ground state with the same electron configuration as the nitrogen atom in a molecule of ammonia. [1]
- 71 Determine the energy required to break all of the bonds in 0.250 mol of H–H bonds in the $\text{H}_2(\text{g})$. [1]
-

Base your answers to questions 72 and 73 on the information below and on your knowledge of chemistry.

The diagram below represents a sealed, rigid cylinder with a movable piston containing helium gas at 296 K and 1.3 atm.



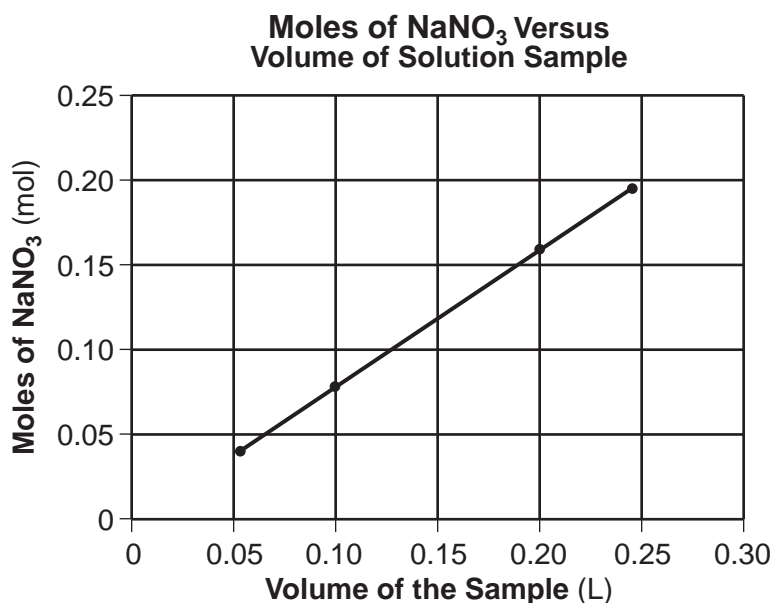
- 72 Compare the density of the helium in the cylinder when the volume is 50.0 mL to the density of the helium in the cylinder when the volume is decreased to 25.0 mL. [1]
- 73 State *one* change in temperature and *one* change in pressure that will cause the gas in the cylinder to behave more like an ideal gas. [1]
-

Base your answers to questions 74 through 76 on the information below and on your knowledge of chemistry.

Four different samples of $\text{NaNO}_3(\text{aq})$ are each evaporated to dryness. The solution volume and mass of the dry $\text{NaNO}_3(\text{s})$ of each sample are recorded in the table below.

Sample	Volume of $\text{NaNO}_3(\text{aq})$ (L)	Mass of dry $\text{NaNO}_3(\text{s})$ (g)
1	0.0524	3.56
2	0.0988	6.72
3	0.2017	13.71
4	0.2431	16.53

The number of moles of $\text{NaNO}_3(\text{s})$ of each sample was then calculated and used to produce the graph below.



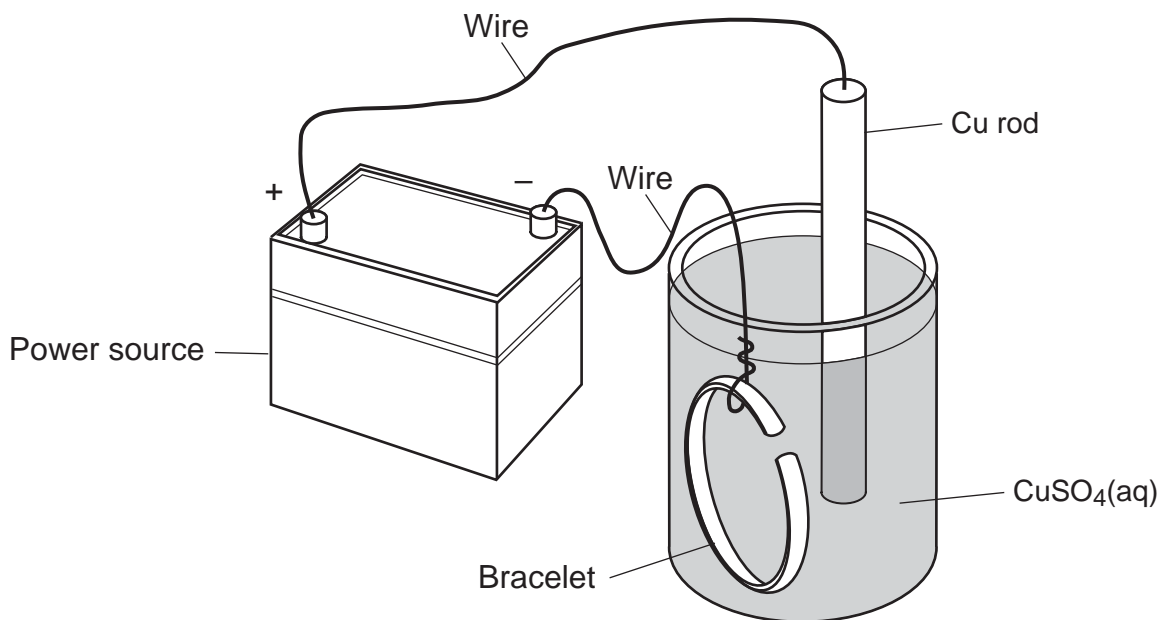
74 Determine the number of moles of NaNO_3 in 0.15 L of the solution. [1]

75 Write a chemical name for NaNO_3 . [1]

76 Compare the boiling point of the solution in sample 1 at standard pressure to the boiling point of water at standard pressure. [1]

Base your answers to questions 77 through 79 on the information below and on your knowledge of chemistry.

Electroplating is an electrolytic process that can be used to coat metal objects with a less reactive metal. The diagram below shows an electroplating cell that includes a power source connected to a copper rod and a bracelet made from a different metal. The rod and bracelet are in an aqueous copper(II) sulfate solution.



77 Identify the electrode that attracts the Cu^{2+} ions as the cell operates. [1]

78 Determine the oxidation state of sulfur in CuSO_4 . [1]

79 Write a balanced half-reaction equation for the reduction of Cu^{2+} ions that occurs in this cell. [1]

Base your answers to questions 80 through 82 on the information below and on your knowledge of chemistry.

In a titration, 15.0 mL of hydrochloric acid, $\text{HCl}(\text{aq})$, of unknown concentration is exactly neutralized by the addition of 8.0 mL of 0.15 M $\text{KOH}(\text{aq})$. The pH value of the $\text{HCl}(\text{aq})$ solution is 1.1 before the titration begins.

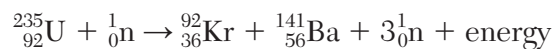
80 Complete the equation *in your answer booklet* for the reaction that occurs in this titration by writing the formula for *each* product. [1]

81 State the pH value of the solution when the $\text{H}^+(\text{aq})$ ion concentration in the acid has *decreased* by a factor of 10 compared to its original value. [1]

82 Determine the molarity of the $\text{HCl}(\text{aq})$, by using the titration data. [1]

Base your answers to questions 83 through 85 on the information below and on your knowledge of chemistry.

Nuclear fission is currently used to produce electricity in a nuclear power plant. One possible fission reaction is represented by the equation below.



The barium-141 decays by beta emission and has a half-life of 18.3 minutes.

- 83 Based on Table N, state the decay mode for U-235. [1]
- 84 Complete the nuclear equation *in your answer booklet* for the decay of Ba-141, by writing a notation for the missing product. [1]
- 85 Compare the amount of energy released by the fission of one mole of uranium-235 to the amount of energy released by the combustion of one mole of octane fuel, C₈H₁₈. [1]
-

P.S./CHEMISTRY

Printed on Recycled Paper

P.S./CHEMISTRY

PHYSICAL SETTING CHEMISTRY

Thursday, August 17, 2023 — 8:30 to 11:30 a.m.

ANSWER BOOKLET

Student

Teacher

School Grade

Record your answers for Part B–2 and Part C in this booklet.

Part B–2

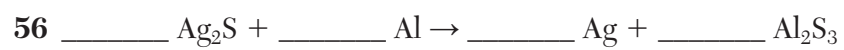
51 _____

52 _____

53 _____

54 _____

55 _____



57 _____ **g/mol**

58

59 _____

60 _____

61 _____

62

63

64

65

Part C

66 _____

67 _____ %

68

69 _____

70 _____

71 _____ kJ

72 _____

73 Temperature: _____

Pressure: _____

74 _____ mol

75 _____

76 _____

77 _____

78 _____

79 _____

80 $\text{HCl(aq)} + \text{KOH(aq)} \rightarrow$ _____ + _____

81 _____

82 _____ M

83 _____

84 ${}^{141}_{56}\text{Ba} \rightarrow {}^0_{-1}\text{e} +$ _____

85 _____

P.S./CHEMISTRY

Printed on Recycled Paper

P.S./CHEMISTRY

Regents Examination in Physical Setting/Chemistry – August 2023**Scoring Key: Parts A and B-1 (Multiple-Choice Questions)**

Examination	Date	Question Number	Scoring Key	Question Type	Credit	Weight
Physical Setting/Chemistry	August '23	1	1	MC	1	1
Physical Setting/Chemistry	August '23	2	2	MC	1	1
Physical Setting/Chemistry	August '23	3	3	MC	1	1
Physical Setting/Chemistry	August '23	4	3	MC	1	1
Physical Setting/Chemistry	August '23	5	2	MC	1	1
Physical Setting/Chemistry	August '23	6	3	MC	1	1
Physical Setting/Chemistry	August '23	7	1	MC	1	1
Physical Setting/Chemistry	August '23	8	3	MC	1	1
Physical Setting/Chemistry	August '23	9	4	MC	1	1
Physical Setting/Chemistry	August '23	10	3	MC	1	1
Physical Setting/Chemistry	August '23	11	4	MC	1	1
Physical Setting/Chemistry	August '23	12	3	MC	1	1
Physical Setting/Chemistry	August '23	13	2	MC	1	1
Physical Setting/Chemistry	August '23	14	2	MC	1	1
Physical Setting/Chemistry	August '23	15	4	MC	1	1
Physical Setting/Chemistry	August '23	16	1	MC	1	1
Physical Setting/Chemistry	August '23	17	4	MC	1	1
Physical Setting/Chemistry	August '23	18	2	MC	1	1
Physical Setting/Chemistry	August '23	19	3	MC	1	1
Physical Setting/Chemistry	August '23	20	3	MC	1	1
Physical Setting/Chemistry	August '23	21	1	MC	1	1
Physical Setting/Chemistry	August '23	22	3	MC	1	1
Physical Setting/Chemistry	August '23	23	1	MC	1	1
Physical Setting/Chemistry	August '23	24	3	MC	1	1
Physical Setting/Chemistry	August '23	25	3	MC	1	1
Physical Setting/Chemistry	August '23	26	2	MC	1	1
Physical Setting/Chemistry	August '23	27	1	MC	1	1
Physical Setting/Chemistry	August '23	28	4	MC	1	1
Physical Setting/Chemistry	August '23	29	3	MC	1	1
Physical Setting/Chemistry	August '23	30	2	MC	1	1
Physical Setting/Chemistry	August '23	31	3	MC	1	1
Physical Setting/Chemistry	August '23	32	4	MC	1	1
Physical Setting/Chemistry	August '23	33	1	MC	1	1
Physical Setting/Chemistry	August '23	34	4	MC	1	1
Physical Setting/Chemistry	August '23	35	2	MC	1	1
Physical Setting/Chemistry	August '23	36	3	MC	1	1
Physical Setting/Chemistry	August '23	37	2	MC	1	1
Physical Setting/Chemistry	August '23	38	2	MC	1	1
Physical Setting/Chemistry	August '23	39	1	MC	1	1
Physical Setting/Chemistry	August '23	40	3	MC	1	1
Physical Setting/Chemistry	August '23	41	2	MC	1	1
Physical Setting/Chemistry	August '23	42	3	MC	1	1
Physical Setting/Chemistry	August '23	43	2	MC	1	1
Physical Setting/Chemistry	August '23	44	2	MC	1	1
Physical Setting/Chemistry	August '23	45	1	MC	1	1
Physical Setting/Chemistry	August '23	46	4	MC	1	1
Physical Setting/Chemistry	August '23	47	4	MC	1	1
Physical Setting/Chemistry	August '23	48	2	MC	1	1
Physical Setting/Chemistry	August '23	49	1	MC	1	1
Physical Setting/Chemistry	August '23	50	3	MC	1	1

Regents Examination in Physical Setting/Chemistry – August 2023

Scoring Key: Parts B-2 and C (Constructed-Response Questions)

Examination	Date	Question Number	Scoring Key	Question Type	Credit	Weight
Physical Setting/Chemistry	August '23	51	-	CR	1	1
Physical Setting/Chemistry	August '23	52	-	CR	1	1
Physical Setting/Chemistry	August '23	53	-	CR	1	1
Physical Setting/Chemistry	August '23	54	-	CR	1	1
Physical Setting/Chemistry	August '23	55	-	CR	1	1
Physical Setting/Chemistry	August '23	56	-	CR	1	1
Physical Setting/Chemistry	August '23	57	-	CR	1	1
Physical Setting/Chemistry	August '23	58	-	CR	1	1
Physical Setting/Chemistry	August '23	59	-	CR	1	1
Physical Setting/Chemistry	August '23	60	-	CR	1	1
Physical Setting/Chemistry	August '23	61	-	CR	1	1
Physical Setting/Chemistry	August '23	62	-	CR	1	1
Physical Setting/Chemistry	August '23	63	-	CR	1	1
Physical Setting/Chemistry	August '23	64	-	CR	1	1
Physical Setting/Chemistry	August '23	65	-	CR	1	1
Physical Setting/Chemistry	August '23	66	-	CR	1	1
Physical Setting/Chemistry	August '23	67	-	CR	1	1
Physical Setting/Chemistry	August '23	68	-	CR	1	1
Physical Setting/Chemistry	August '23	69	-	CR	1	1
Physical Setting/Chemistry	August '23	70	-	CR	1	1
Physical Setting/Chemistry	August '23	71	-	CR	1	1
Physical Setting/Chemistry	August '23	72	-	CR	1	1
Physical Setting/Chemistry	August '23	73	-	CR	1	1
Physical Setting/Chemistry	August '23	74	-	CR	1	1
Physical Setting/Chemistry	August '23	75	-	CR	1	1
Physical Setting/Chemistry	August '23	76	-	CR	1	1
Physical Setting/Chemistry	August '23	77	-	CR	1	1
Physical Setting/Chemistry	August '23	78	-	CR	1	1
Physical Setting/Chemistry	August '23	79	-	CR	1	1
Physical Setting/Chemistry	August '23	80	-	CR	1	1
Physical Setting/Chemistry	August '23	81	-	CR	1	1
Physical Setting/Chemistry	August '23	82	-	CR	1	1
Physical Setting/Chemistry	August '23	83	-	CR	1	1
Physical Setting/Chemistry	August '23	84	-	CR	1	1
Physical Setting/Chemistry	August '23	85	-	CR	1	1

Key
MC = Multiple-choice question
CR = Constructed-response question

The chart for determining students' final examination scores for the **August 2023 Regents Examination in Physical Setting/Chemistry** will be posted on the Department's web site at <https://www.nysedregents.org/Chemistry/> on the day of the examination. Conversion charts provided for the previous administrations of the Physical Setting/Chemistry examination must NOT be used to determine students' final scores for this administration.

FOR TEACHERS ONLY

The University of the State of New York
REGENTS HIGH SCHOOL EXAMINATION

PHYSICAL SETTING/CHEMISTRY

Thursday, August 17, 2023 — 8:30 to 11:30 a.m., only

RATING GUIDE

Directions to the Teacher:

Refer to the directions on page 2 before rating student papers.

Updated information regarding the rating of this examination may be posted on the New York State Education Department's web site during the rating period. Check this web site at: <https://www.nysed.gov/state-assessment/high-school-regents-examinations> and select the link "Scoring Information" for any recently posted information regarding this examination. This site should be checked before the rating process for this examination begins and several times throughout the Regents Examination period.

Directions to the Teacher

Follow the procedures below for scoring student answer papers for the Regents Examination in Physical Setting/Chemistry. Additional information about scoring is provided in the publication *Information Booklet for Scoring Regents Examinations in the Sciences*.

At least two science teachers must participate in the scoring of the Part B–2 and Part C open-ended questions on a student’s paper. Each of these teachers should be responsible for scoring a selected number of the open-ended questions on each answer paper. No one teacher is to score more than approximately one-half of the open-ended questions on a student’s answer paper. Teachers may not score their own students’ answer papers.

Students’ responses must be scored strictly according to the Rating Guide. For open-ended questions, credit may be allowed for responses other than those given in the rating guide if the response is a scientifically accurate answer to the question and demonstrates adequate knowledge, as indicated by the examples in the rating guide. Do not attempt to correct the student’s work by making insertions or changes of any kind. On the student’s separate answer sheet, for each question, record the number of credits earned and the teacher’s assigned rater/scorer letter.

Fractional credit is *not* allowed. Only whole-number credit may be given for a response. If the student gives more than one answer to a question, only the first answer should be rated. Units need not be given when the wording of the questions allows such omissions.

For hand scoring, raters should enter the scores earned in the appropriate boxes printed on the separate answer sheet. Next, the rater should add these scores and enter the total in the box labeled “Total Raw Score.” Then the student’s raw score should be converted to a scale score by using the conversion chart that will be posted on the Department’s web site at: <https://www.nysed.gov/state-assessment/high-school-regents-examinations> on Thursday, August 17, 2023. The student’s scale score should be entered in the box labeled “Scale Score” on the student’s answer sheet. The scale score is the student’s final examination score.

Schools are not permitted to rescore any of the open-ended questions on this exam after each question has been rated once, regardless of the final exam score. Schools are required to ensure that the raw scores have been added correctly and that the resulting scale score has been determined accurately.

Because scale scores corresponding to raw scores in the conversion chart may change from one administration to another, it is crucial that, for each administration, the conversion chart provided for that administration be used to determine the student’s final score.

Part B–2

Allow a total of 15 credits for this part. The student must answer all questions in this part.

51 [1] Allow 1 credit for 6 *or* six.

52 [1] Allow 1 credit for 17 *or* seventeen.

53 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

Metals conduct electricity better than halogens.

A halogen is a poor conductor compared to an alkali metal.

The metal is better.

54 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

ionic bonds

ionic

55 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

All halogen atoms have 7 valence electrons in the ground state.

They have atoms with the same number of outermost shell electrons.

They each need one more electron to achieve a stable octet.

56 [1] Allow 1 credit for $\underline{3}$ Ag₂S + $\underline{2}$ Al → $\underline{6}$ Ag + $\underline{\quad}$ Al₂S₃.

Allow 1 credit even if the coefficient “1” is written in front of Al₂S₃.

57 [1] Allow 1 credit for 150. g/mol *or* for any value from 149.96 g/mol to 150.3 g/mol, inclusive.

58 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

$$\text{moles} = \frac{546 \text{ g}}{248 \text{ g/mol}}$$

$$546 \text{ g} \times \frac{1 \text{ mol}}{248 \text{ g}}$$

$$\frac{1}{248} = \frac{x}{546}$$

$$\frac{546}{248}$$

59 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

The equation shows energy on the reactant side.

Energy is on the left side of the equation.

The 53 kJ is on the left side.

heat term on reactant side

60 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

At equilibrium, the forward reaction rate equals the reverse reaction rate, so equal numbers of HI molecules are produced and broken down.

The rate of formation of HI equals the rate at which HI is decomposed.

The rates are equal.

61 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

An increase in temperature will increase the concentration of HI(g).

The concentration of HI(g) increases.

HI increases.

62 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

Butane has only single carbon-carbon bonds.

There are no multiple bonds between the carbon atoms.

no C=C or C≡C bonds; just C–C bonds

63 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

halide

halocarbon

alkyl halide

halogenoalkanes

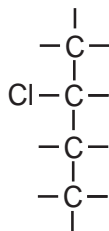
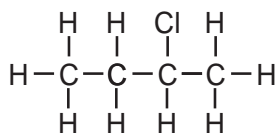
64 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

The intermolecular forces in 1-chlorobutane are stronger than the intermolecular forces in 2-chlorobutane.

weaker in 2-chlorobutane

65 [1] Allow 1 credit.

Examples of 1-credit responses:



Note: Do *not* allow credit if only some of the H atoms bonded to C atoms are shown.

Part C

Allow a total of 20 credits for this part. The student must answer all questions in this part.

66 [1] Allow 1 credit for 3 *or* three.

67 [1] Allow 1 credit for 34.0% *or* 34%.

68 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

$$\frac{37.0\% - 36.1\%}{36.1\%} \times 100 =$$

$$\frac{37.0 - 36.1}{36.1} \times 100 =$$

$$\frac{0.9\%}{36.1\%} \times 100 =$$

$$\frac{90}{36.1}$$

Note: Do *not* allow credit if the fraction is not multiplied by 100.

69 [1] Allow 1 credit for 2 *or* two *or* 1 pair.

70 [1] Allow 1 credit for Ne *or* neon.

71 [1] Allow 1 credit for 109 kJ.

72 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

The density of the helium is less at 50.0 mL.

The He is more dense at 25.0 mL.

less dense at 50 mL

73 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

Temperature: higher

Pressure: lower

Temperature: any temperature above 296 K

Pressure: any pressure lower than 1.3 atm

74 [1] Allow 1 credit for any value from 0.11 mol to 0.13 mol, inclusive.

75 [1] Allow 1 credit for sodium nitrate.

76 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

The sample 1 solution has a higher boiling point than water, when both are at standard pressure.

Water has a lower boiling point than the solution.

The BP of $\text{NaNO}_3(\text{aq})$ is higher.

lower for water

77 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

The Cu^{2+} ions are attracted to the bracelet.

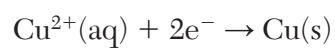
bracelet

negative electrode

cathode

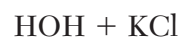
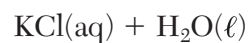
78 [1] Allow 1 credit for +6 *or* 6+.

79 [1] Allow 1 credit. Acceptable responses include, but are not limited to:



Note: Do *not* allow credit for the e without the minus sign (-).

80 [1] Allow 1 credit. Acceptable responses include, but are not limited to:



81 [1] Allow 1 credit for 2.1.

82 [1] Allow 1 credit for 0.080 M *or* 0.08 M.

83 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

alpha decay

alpha

α

${}^4_2\text{He}$

${}^4_2\alpha$

alpha particle

84 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

${}^{141}_{57}\text{La}$

${}^{141}\text{La}$

La-141

lanthanum-141

85 [1] Allow 1 credit. Acceptable responses include, but are not limited to:

The fission of one mole of U-235 releases much more energy than the combustion of one mole of C_8H_{18} .

For equal quantities, fission gives out more energy than combustion.

Burning 1 mol of C_8H_{18} releases much less energy.

less energy from chemical reaction

Regents Examination in Physical Setting/Chemistry

August 2023

Chart for Converting Total Test Raw Scores to Final Examination Scores (Scale Scores)

The *Chart for Determining the Final Examination Score for the August 2023 Regents Examination in Physical Setting/Chemistry* will be posted on the Department's web site at: <https://www.nysed.gov/state-assessment/high-school-regents-examinations> on Thursday, August 17, 2023. Conversion charts provided for previous administrations of the Regents Examination in Physical Setting/Chemistry must NOT be used to determine students' final scores for this administration.

Online Submission of Teacher Evaluations of the Test to the Department

Suggestions and feedback from teachers provide an important contribution to the test development process. The Department provides an online evaluation form for State assessments. It contains spaces for teachers to respond to several specific questions and to make suggestions. Instructions for completing the evaluation form are as follows:

1. Go to <https://www.surveymonkey.com/r/8LNLLDW>.
2. Select the test title.
3. Complete the required demographic fields.
4. Complete each evaluation question and provide comments in the space provided.
5. Click the SUBMIT button at the bottom of the page to submit the completed form.

Map to Core Curriculum

August 2023 Physical Setting/Chemistry			
Question Numbers			
Key Ideas/Performance Indicators	Part A	Part B	Part C
Standard 1			
Math Key Idea 1		32, 38, 58	66, 68, 74
Math Key Idea 2		43, 54	79
Math Key Idea 3		36, 41, 52, 56, 57	67, 71, 78, 82
Science Inquiry Key Idea 1		33, 34, 40, 42, 44, 46, 47, 48, 49, 51, 54, 60, 61, 62, 63, 64	69, 70, 72, 73, 76, 83, 85
Science Inquiry Key Idea 2			
Science Inquiry Key Idea 3		31, 34, 40, 42, 46, 47, 48, 49, 50, 56, 59, 64	69, 70, 71, 73, 75, 77, 78, 79, 80, 84
Engineering Design Key Idea 1			
Standard 2			
Key Idea 1			
Key Idea 2		50	
Key Idea 3			
Standard 6			
Key Idea 1			
Key Idea 2		65	68
Key Idea 3			81
Key Idea 4			
Key Idea 5			
Standard 7			
Key Idea 1			
Key Idea 2			
Standard 4 Process Skills			
Key Idea 3		31, 32, 33, 35, 36, 39, 45, 47, 49, 51, 52, 53, 55, 56, 57, 58, 60, 61, 63, 65	77, 79, 80, 82
Key Idea 4		41, 43, 59	84
Key idea 5		37	70
Standard 4			
Key Idea 3	1, 2, 3, 4, 5, 6, 7, 8, 10, 11, 14, 15, 16, 17, 18, 19, 22, 23, 24, 25, 26, 27, 29, 30	31, 32, 33, 34, 35, 36, 39, 40, 42, 44, 45, 46, 47, 48, 49, 51, 52, 53, 55, 56, 57, 58, 60, 61, 62, 63, 65	66, 67, 68, 72, 73, 74, 75, 76, 77, 78, 79, 80, 81, 82
Key Idea 4	21	41, 43, 50, 59	83, 84
Key Idea 5	9, 12, 13, 20, 28	37, 38, 54, 64	69, 70, 71, 85
Reference Tables			
2011 Edition	3, 4, 5, 6, 7, 13, 14, 16, 24, 25, 27, 30	31, 34, 35, 37, 38, 40, 41, 42, 44, 47, 49, 50, 51, 52, 55, 57, 58, 59, 62, 63, 65	67, 68, 70, 71, 72, 75, 78, 82, 83, 84, 85

Regents Examination in Physical Setting/Chemistry – August 2023

Chart for Converting Total Test Raw Scores to Final Examination Scores (Scale Scores)

Raw Score	Scale Score	Raw Score	Scale Score	Raw Score	Scale Score	Raw Score	Scale Score
85	100	63	73	41	59	19	39
84	98	62	73	40	58	18	38
83	96	61	72	39	57	17	36
82	94	60	71	38	57	16	35
81	93	59	70	37	56	15	34
80	91	58	70	36	55	14	32
79	90	57	69	35	54	13	31
78	88	56	68	34	54	12	29
77	87	55	68	33	53	11	27
76	86	54	67	32	52	10	25
75	85	53	66	31	51	9	24
74	83	52	66	30	50	8	22
73	82	51	65	29	49	7	19
72	81	50	64	28	49	6	17
71	80	49	64	27	48	5	15
70	79	48	63	26	47	4	12
69	78	47	63	25	46	3	10
68	77	46	62	24	45	2	7
67	77	45	61	23	44	1	4
66	76	44	61	22	43	0	0
65	75	43	60	21	41		
64	74	42	59	20	40		

To determine the student's final examination score, find the student's total test raw score in the column labeled "Raw Score" and then locate the scale score that corresponds to that raw score. The scale score is the student's final examination score. Enter this score in the space labeled "Scale Score" on the student's answer sheet.

Schools are not permitted to rescore any of the open-ended questions on this exam after each question has been rated once, regardless of the final exam score. Schools are required to ensure that the raw scores have been added correctly and that the resulting scale score has been determined accurately.

Because scale scores corresponding to raw scores in the conversion chart change from one administration to another, it is crucial that for each administration the conversion chart provided for that administration be used to determine the student's final score. The chart above is usable only for this administration of the Regents Examination in Physical Setting/Chemistry.